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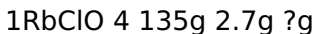
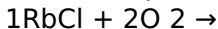
Chapter 12 Practice
Test Stoichiometry
Form A Sample: Given
the equation $\text{RbCl} + \text{O}$

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2 RbClO₄, find the
mass of RbClO₄
formed if 2.7g of O₂
reacts with 135g RbCl.

Analyze, rewrite,
balance equation



135g 2.7g ?g

Solve for 1st reactant

135g RbCl x 1 mol

RbCl x 1 mol RbClO₄
184.917g RbClO₄

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Chapter 12 Multiple
Choice Identify the
letter of the choice that
best completes the

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statement or answers
the question. ____ 1.

The first step in most
stoichiometry problems
is to ____.

- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given quantities to masses

____ 2.

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example problem 1.

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e. 12 g 7. How many
grams of nitric acid,
 HNO_3 , can be
prepared from the
reaction of 138 g of NO_2
with 54.0 g H_2O
according to the

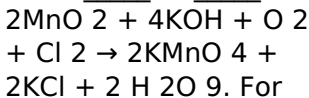
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equation below? 3NO_2
 $+ \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$...

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Name _____ Per _____



9. For the reaction above, there is 100. g of each reactant

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MULTIPLE CHOICE.

Choose the one alternative that best completes the statement or answers the question. 1)

Consider the following reaction: $3A \rightarrow 2B$ The average rate of appearance of B is given by $D[B]/Dt$.

Comparing the rate of appearance of B and the rate of

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MULTIPLE ...

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Practice Problems

(Chapter 5):

Stoichiometry CHEM

30A Part I: Using the
conversion factors in
your tool box g A mol A

mol A 1. How many
moles CH₃OH are in
14.8 g CH₃OH? 2.

What is the mass in
grams of 1.5×10^{16}
atoms S? 3. How many
molecules of CO₂ are

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in 12.0 g CO₂? 2 4.

What is the mass in grams of 1 atom of Au?

KEY Tool Box: To ...

Practice Problems

(Chapter 5):

Stoichiometry

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Stoichiometry Test B ...

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Stoichiometry Name

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Stoichiometry (page 2 of 2) 19 The mass of element X found in 100 mole of each of four different compounds is 280 g, 420 g, 560 g,

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Stoichiometry Test B
- Reliefwatch

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the first step is always

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